**UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS**

**GCE Advanced Level**

# MARK SCHEME for the October/November 2011 question paper for the guidance of teachers

## 9701 CHEMISTRY

**9701/41** Paper 4 (A2 Structured Questions), maximum raw mark 100

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

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**1 (a) (i)** *either* burn *or* shine light/uv on mixture of $H_2$ + $Cl_2$ *but* NOT heat [1]

**(ii)** red/orange/brown colour of bromine decolourises/disappears
steamy/misty/white fumes produced
container gets warm/hot [2]

**(iii)** H-H = 436 Cl-Cl = 244 H-Cl = 431

$\Delta H$ = 436 + 244 – 2(431) = –182 kJ $mol^{-1}$ [2]

H-H = 436 Br-Br = 193 H-Br = 366

$\Delta H$ = 436 + 193 – 2(366) = –103 kJ $mol^{-1}$ [2]

**(iv)** H-Br bond is weaker than the H-Cl bond – allow converse. [1]
**[8]**

**(b) (i)** light [1]

**(ii)** bonds broken = C-H & I-I = 410 + 151 = 561
bonds made = C-I & H-I = 240 + 299 = 539
$\Delta H$ = 551 – 539 = +22 kJ $mol^{-1}$ [2]

**(iii)** The overall reaction is endothermic *or* no strong bonds/only weak bonds are formed *or* high $E_{act}$ [1]
**[4]**

**(c) (i)** homolytic fission is the breaking of a bond to form (two) radicals/neutral species/ odd-electron species [1]

**(ii)** $\bullet CH_2Cl$ [1]
the C-Br bond is the weakest or needs least energy to break/breaks most easily [1]
**[3]**

**(d)**


4 structures: [2]
2 or 3 structures: [1]

Correct chiral atom identified [1]
**[3]**

**[Total: 18]**

**2 (a) (i)** Order w.r.t. $[CH_3CHO] = 1$ [1]
Order w.r.t. $[CH_3OH] = 1$ [1]
Order w.r.t. $[H^+] = 1$ [1]

**(ii)** rate = $k[CH_3CHO][CH_3OH][H^+]$ [1]

**(iii)** units = $mol^{-2}\ dm^6\ s^{-1}$ [1]

**(iv)** rate will be 2 × 4 = 8 times as fast as reaction 1 (relative rate = 8) [1]
**[6]**

**(b)**

| | $[CH_3CHO]$ /mol $dm^{-3}$ | $[CH_3OH]$ /mol $dm^{-3}$ | $[H^+]$ /mol $dm^{-3}$ | [acetal **A**] /mol $dm^{-3}$ | $[H_2O]$ /mol $dm^{-3}$ |
|---|---|---|---|---|---|
| at start | 0.20 | 0.10 | 0.05 | 0.00 | 0.00 |
| at equilibrium | (0.20 – x) | ***(0.10 – 2x)*** | ***0.05*** | x | ***x*** |
| at equilibrium | ***0.175*** | ***0.05*** | ***0.05*** | 0.025 | ***0.025*** |

**(i)** 3 values in second row 3 x [1]

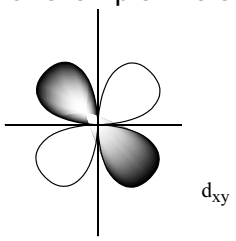
**(ii)** 4 values in third row 4 x [1]

**(iii)** $K_c = \{[\text{acetal } \mathbf{A}][H_2O]\}/\{[CH_3CHO][CH_3OH]^2\}$ [1]
units = $mol^{-1}dm^3$ [1]

**(iv)** $K_c = 0.025^2/(0.175 \times 0.05^2) =$ **1.4(3)** ($mol^{-1}\ dm^3$) [1]
**[max 9]**

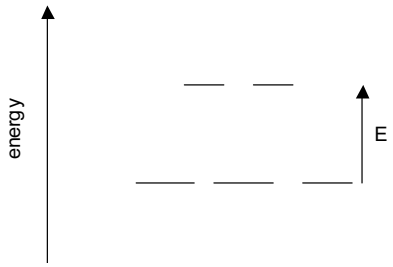
**[Total: 15]**

**3** **(a)** for example.... also allow $d_{z2}$




shape (4 lobes) [1]
correct label e.g. $d_{xy}$ [1]
**[2]**

**(b)** **(i)**




Marks are for 5 degenerate orbitals [1]
and 3:2 split [1]

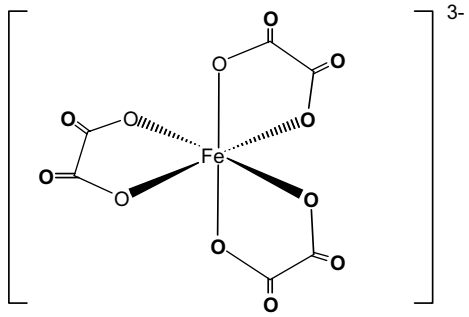
**(ii)** colour due to the absorption of light NOT emitted light [1]
E = hf *or* photon's energy = E in above diagram [1]
electron promoted from lower to higher orbital [1]

size of ΔE depends on the ligand [1]
as ΔE changes, so does f in E = hf [1]
**[7]**

**(c)** **(i)** O.N.(carbon) = +3 (4 × (–2) + 2x = –2, thus 2x = +6) [1]

**(ii)** O.N. = +3 [1]

**(iii)**




[2]

**(iv)** **2** $K_3Fe(C_2O_4)_3 \rightarrow$ **3** $K_2C_2O_4$ + **2** $FeC_2O_4$ + **2** $CO_2$ [2]
*Or* $K_3Fe(C_2O_4)_3 \rightarrow$ **3/2** $K_2C_2O_4 + FeC_2O_4 + CO_2$

**[max 5]**

**[Total: 14]**

**4 (a) (i)** $C_2H_5NH_2 + HA \rightarrow C_2H_5NH_3^+ + A^-$ (HA can be $H_2O$, HC*l* etc.) [1]
Allow $\rightleftharpoons$ instead of arrow

**(ii)**

| most basic | | least basic |
|---|---|---|
| **ethylamine** | **ammonia** | **phenylamine** |

[1]

**(iii)** ethylamine > $NH_3$ due to electron-donating ethyl/alkyl group [1]
phenylamine < $NH_3$ due to delocalisation of lone pair over ring [1]
**[4]**

**(b) (i)** $C_6H_5OH + OH^- \rightarrow C_6H_5O^- + H_2O$ (*or* with $Na^+/H_2O/A^-$) [1]

**(ii)** pKa of nitrophenol is smaller/$K_a$ is larger because it's a stronger acid/dissociates more than phenol [1]
stronger because the anionic charge is spread out moreover the $NO_2$ group *or* $NO_2$ is electron-withdrawing [1]

**(iii)** pKa = 1.0 [1]

**(iv)** Nitro group increases acidity / electron-withdrawing groups increase acidity [1]
**[5]**

**(c) (i)** **B** is phenyldiazonium cation, $C_6H_5–N^+\equiv N$ [1]

**(ii)**

| reaction | reagent(s) | conditions |
|---|---|---|
| Step 1 | **$NaNO_2$ + HC*l*<br>or $HNO_2$** [1] | **T < 10°C** [1] |
| Step 2 | **$H_2O$ / aq** | **heat/boil/T > 10°** (both) [1] |
| Step 3 | **$HNO_3$<br>NB $HNO_3$(aq) OK for both** | **dilute** (both) [1] |

[4]
**[5]**

**[Total: 14]**

**5 (a) (i)** C=C double bonds / alkenes

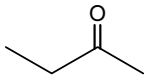
**(ii)** –OH groups / accept alcohols or acids

**(iii)** $CH_3CO–$ or $CH_3CH(OH)–$ groups

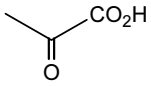
**(iv)** carbonyl, >C=O, groups / accept aldehydes and ketones 4 × [1]

**[4]**

**(b)**



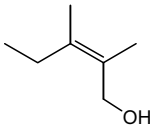
**D**



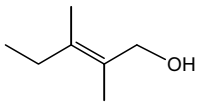
**E** 2 × [1]

**[2]**

**(c)** isomers of **C**



cis



trans

correct structure (excl. stereochemistry) [1]
cis and trans drawn correctly [1]
type of isomerism is **cis-trans *or* geometrical isomerism** [1]

**[3]**

**[Total: 9]**

**6 (a) (i)** $2H_2NCH_2CO_2H \rightarrow H_2NCH_2CONHCH_2CO_2H + H_2O$ [1]

**(ii)** Skeletal formula required [1]
**[2]**

**(b) (i)** α helix [1]
β pleated sheet [1]

**(ii) Students should choose one of the structures below**

| For α helix: | For β pleated sheet: |
|---|---|
| Need to show a helix with C=O - - - H-N between turns | Need to show two parallel 'zig-zag' strands with C=O - - - H-N between them |

Whichever is chosen, overall structure [1] position of H bonds [1]

**[4]**

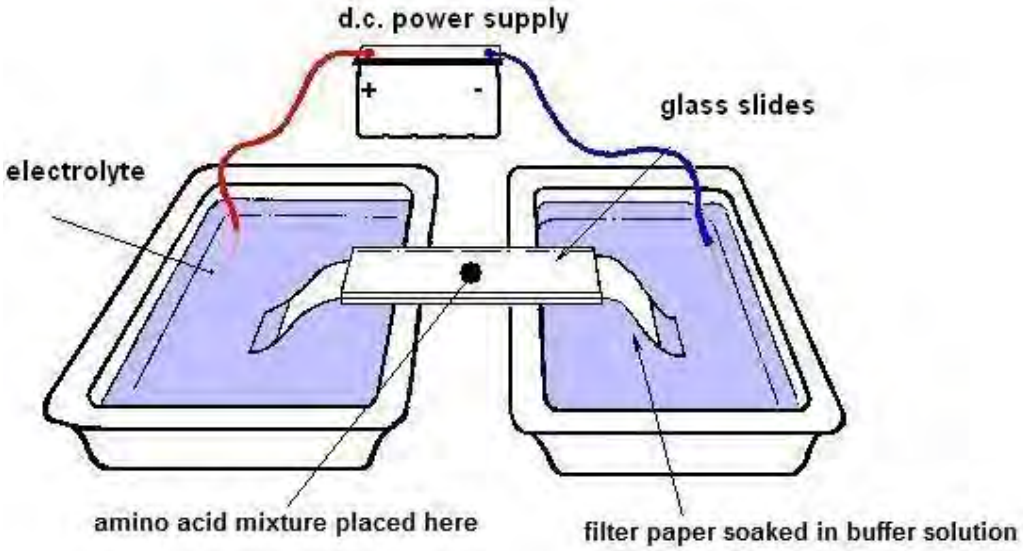
**(c)**

| amino acid residue 1 | amino acid residue 2 | type of bonding |
|---|---|---|
| $–HNCH(CH_2CH_2CH_2CH_2NH_2)CO–$ | $HNCH(CH_2CH_2CO_2H)CO–$ | Ionic bonds or hydrogen bonds |
| $–HNCH(CH_3)CO–$ | $–HNCH(CH_3)CO–$ | van der Waals' |
| $–HNCH(CH_2SH)CO–$ | $–HNCH(CH_2SH)CO–$ | Disulfide bonds |
| $–HNCH(CH_2OH)CO–$ | $–HNCH(CH_2CO_2H)CO–$ | Hydrogen bonds |

**[4]**

**[Total: 10]**

**7** **(a)** Sketch and label the apparatus used to carry out electrophoresis e.g




Marks: power supply / electrolyte + filter paper / buffer / acid mixture central — 4 × [1]
**[4]**

**(b)** **(i)** pH of the buffer — [1]
Charge on the amino acid species — [1]

**(ii)** Size of the amino acid species / $M_r$ — [1]
Voltage applied — [1]
Magnitude of the charge (on the amino acid species) — [1]
Temperature — [1]
(max 3)
**[max 3]**

**(c)** **(i)** They have insufficient electron density / only one electron — [1]
**(ii)** Sulfur — [1]
because it has the greatest atomic number / number of electrons — [1]
**[3]**

**[Total: 10]**

**8 (a)**

| traditional material | modern polymer used |
|---|---|
| Paper/cardboard/wood/leaves<br>hessian/hemp/jute<br>steel/aluminium | PVC in packaging |
| Cotton/wool/linen | *Terylene* in fabrics |
| Glass/china/porcelain/earthenware<br>metal/leather | Polycarbonate bottle |

3 → 2 marks, 2 → 1 mark
**[2]**

**(b)** Reasons: Plastics/polymers pollute the environment for a long time do not decompose/ biodegrade quickly [1]
They are mainly produced from oil [1]
Produce toxic gases on burning [1]
max two

Strategy 1: Recycle polymer waste / use renewable resources [1]
Strategy 2: Develop biodegradable polymers [1]
**[max 3]**

**(c)** PVC [1]
Combustion would produce HC*l* / dioxins as a pollutant [1]
**or**
nylon/acrylic [1]
Combustion would produce HCN [1]
**[2]**

**(d) (i)** Polythene (or other addition polymer) [1]

**(ii)** Addition polymerisation [1]

The polymer chains don't have strong bonds between them – easy to melt [1]
Could be answered with a suitable diagram **[3]**

**[Total: 10]**