

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**  
GCE Advanced Subsidiary Level and GCE Advanced Level

## **MARK SCHEME for the May/June 2013 series**

### **9701 CHEMISTRY**

**9701/23**

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

Page 2	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2013	9701	23

1 (a) (i)



S atom has 6 **and** C atom has 4 electrons (1)

S=C double bonds (4 electrons) clearly shown (1)

(ii) linear **and** 180° (1) [3]

(b) (i)  $\text{CS}_2 + 3\text{O}_2 \rightarrow \text{CO}_2 + 2\text{SO}_2$  (1)

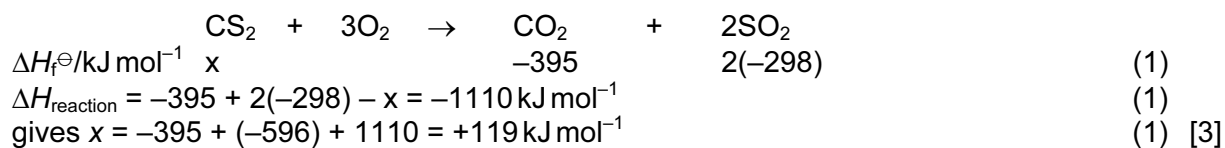
(ii) enthalpy change when 1 mol of a substance (1)

is burnt in an excess of oxygen/air

**or** is completely combusted

under standard conditions (1) [3]

(c)



(d) (i)  $\text{CS}_2 + 2\text{NO} \rightarrow \text{CO}_2 + 2\text{S} + \text{N}_2$   
**or**  
 $\text{CS}_2 + 2\text{NO} \rightarrow \text{CO} + 2\text{S} + \text{N}_2\text{O}$

correct products (1)

correct equation (1)

(ii) from -2 to 0 **both** required (1) [3]

**[Total: 12]**

<b>Page 3</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>GCE AS/A LEVEL – May/June 2013</b>	<b>9701</b>	<b>23</b>

- 2 (a) (i) if the conditions of a system in equilibrium are changed (1)  
the position of equilibrium moves so as to reduce that change (1) [2]
- (ii) lower temperature (1)  
because the forward reaction is exothermic (1)  
higher pressure (1)  
because the forward reaction shows a reduction in volume  
**or**  
there are fewer molecules/moles on RHS of equilibrium (1) [4]

(b)

	CO <sub>2</sub>	+	H <sub>2</sub>	⇌	CO	+	H <sub>2</sub> O	
initial moles	0.70		0.70		0.30		0.30	
equil. moles	(0.70-x)		(0.70-x)		(0.30+x)		(0.30+x)	(1)
equil. concn.	$\frac{(0.70-x)}{1}$		$\frac{(0.70-x)}{1}$		$\frac{(0.30+x)}{1}$		$\frac{(0.30+x)}{1}$	

$$K_c = \frac{(0.30+x)^2}{(0.70-x)^2} = 1.44 \quad (1)$$

gives  $x = 0.25$  (1)

at equilibrium,

$$n(\text{CO}_2) = n(\text{H}_2) = 0.70 - 0.25 = 0.45 \text{ moles}$$

**and**

$$n(\text{CO}) = n(\text{H}_2\text{O}) = 0.3 + 0.25 = 0.55 \text{ moles} \quad (1) \quad [4]$$

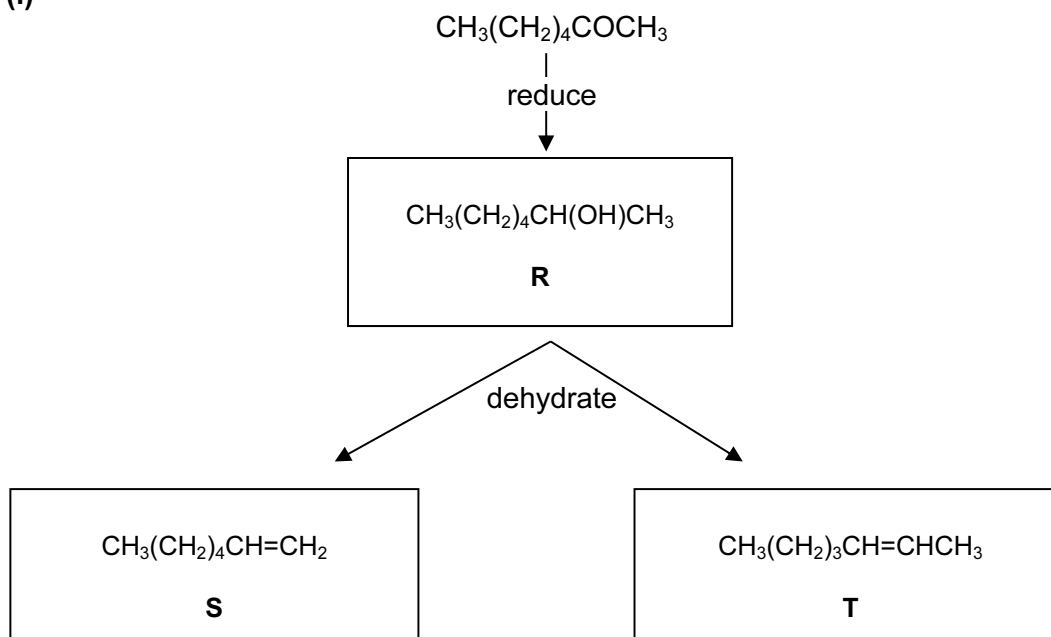
**[Total: 10]**

Page 4	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2013	9701	23

- 3 (a) (i) He or Ne or Ar or Kr (1)
- (ii) P or As (1)
- (iii) Br (1)
- (iv) Na allow Ar (1)
- (v) Si (1)
- (vi) P allow Si (1)
- (vii) Cl or F or Br (1) [7]
- (b) (i) any **two** from  $P_4O_6$ ,  $SO_2$  and  $Cl_2O_7$  (1+1)
- (ii)  $Al_2O_3$  or  $SiO_2$  (1)
- (iii)  $MgSO_3$  (1) [4]
- (c) (i) Si is giant molecular/giant covalent **or**  
P, S, and Cl are simple molecular (1)
- (ii) the molecules are  $S_8$ ,  $P_4$ ,  $Cl_2$  (1)
- larger molecules have more electrons (1)
- and hence greater van der Waals' forces (1) [4]

**[Total: 15]**

4 (a) (i)



one mark for each correct compound, **R**, **S** and **T**

allow correct *cis* and *trans* versions of compound **T** for 2 marks (3 × 1)

(ii) reduction

$\text{NaBH}_4$  or  $\text{LiAlH}_4$  or  $\text{H}_2/\text{Ni}$  or  $\text{Na}/\text{C}_2\text{H}_5\text{OH}$  (1)

dehydration

$\text{P}_4\text{O}_{10}/\text{P}_2\text{O}_5$  or  $\text{H}_3\text{PO}_4$  or conc.  $\text{H}_2\text{SO}_4$  or  $\text{Al}_2\text{O}_3$  (1) [5]

(b)

Tollens' reagent	NO REACTION
HCN	$  \begin{array}{c}  \text{CH}_3(\text{CH}_2)_4\text{C}(\text{OH})\text{CH}_3 \\    \\  \text{CN}  \end{array}  $
$\text{K}_2\text{Cr}_2\text{O}_7/\text{H}^+$	NO REACTION

one mark for each correct answer (3 × 1) [3]

<b>Page 6</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>GCE AS/A LEVEL – May/June 2013</b>	<b>9701</b>	<b>23</b>

(c)  $\text{Na}_2\text{CO}_3$  or  $\text{NaHCO}_3$  effervescence/colourless gas

or

Na colourless gas

or

$\text{PCl}_3/\text{PCl}_5$  etc. steamy fumes

or

$\text{C}_2\text{H}_5\text{OH}/\text{conc. H}_2\text{SO}_4$  sweet smell of ester

or

$\text{K}_2\text{Cr}_2\text{O}_7/\text{H}^+$  orange solution becomes green

correct reagent

(1)

correct observation

(1) [2]

**[Total: 10]**

Page 7	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2013	9701	23

- 5 (a) (i)  $\text{CH}_2=\text{CHCO}_2\text{H}$  (1)
- (ii)  $\text{BrCH}_2\text{CHBrCH}_2\text{OH}$  (1)
- (iii) product is  $\text{HOCH}_2\text{CH}(\text{OH})\text{CH}_2\text{OH}$   
correct addition across  $>\text{C}=\text{C}<$  (1)  
original  $-\text{CH}_2\text{OH}$  remains (1)
- (iv)  $\text{HO}_2\text{CCO}_2\text{H}$  (1) [5]
- (b) (i) nucleophilic substitution (1)
- (ii) oxidation (1) [2]
- (c) (i) **step I**  
 $\text{H}_2$  (1)  
heat with Ni catalyst (1)
- step II**  
acidified  $\text{K}_2\text{Cr}_2\text{O}_7$  (1)  
heat **or** distil off product (1)
- (ii) structural isomerism  
**or**  
functional group isomerism (1) [5]
- (d) **both** oxidation **and** reduction have occurred **or**  
disproportionation has taken place (1) [1]

[Total: 13]